Chemistry 11 - Unit 1 Condensed Review

Cation = Positive (+) Ion

Anion = Negative (-) Ion

PREFIX

Mono = 1
Di = 2
Tri = 3
Tetra = 4
Penta = 5
Hexa = 6
Hepta = 7
Octa = 8
Nona = 9
Deca = 10

IONIC COMPOUND NAMES

Ionic Compounds = Metal + Non-metal

Metal Name + Non-Metal Name (change ending to "-ide")

E.g: Calcium Chlor**ide** **($CaCl_2$)**

Copper (II) Iod**ide** **(CuI_2)**

COVALENT COMPOUND NAMES

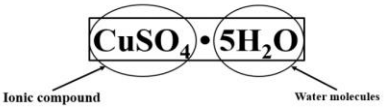
Covalent Compounds = **Non-Metal** + **Non-Metal**

Prefix+Non-Metal Prefix+Non-Metal("-ide")

E.g: **Tri**nitrogen **Deca**brom**ide** **(N_3Br_{10})**

HYDRATES

Ionic Compound + **Prefix**hydrate



E.g: Copper (II) Sulfate **Penta**hydrate
(Ionic Name) **(Prefix Hydrate)**

ACID NAMES

No Oxygen: **"Hydro_____ic Acid"** (Ex. Hydrobromic Acid – HBr)

With Oxygen: Identify the polyatomic oxyanion, and

- Change the endings: **"ite" → "ous"**
  "ate" → "ic"

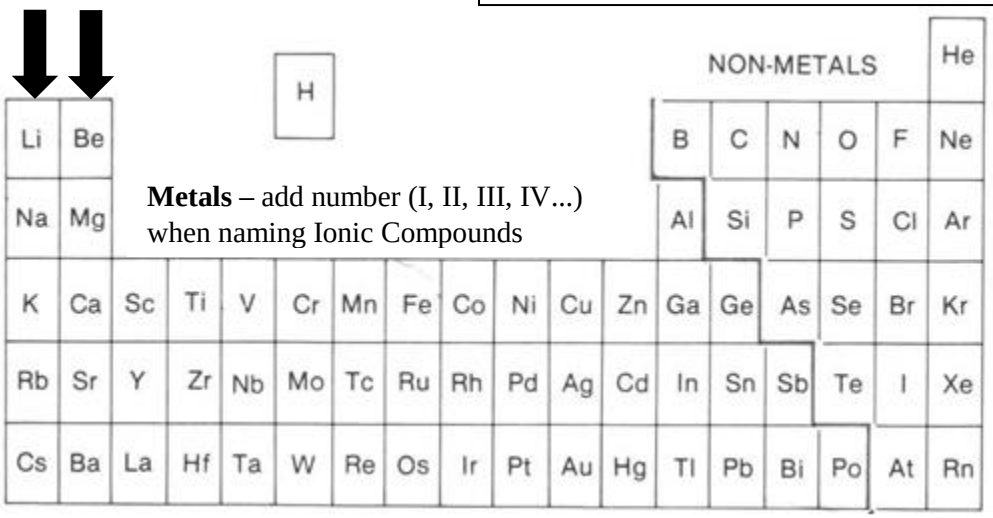
Name the oxoanion and add the word "acid".
(Example: Chlorous Acid – $HClO_2$)

BASE NAMES

1. Name the metal first
2. The polyatomic ion "hydroxide" (-OH) after

Example: NaOH - **Sodium Hydroxide**

Metals
(Do not need to add number when naming Ionic Compounds)



Metals – add number (I, II, III, IV...) when naming Ionic Compounds